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12. Propose a possible way to calculate the average atomic mass of 100 magnesium atoms. Your answer may include a mathematical equation, but it is not required. Add up the masses of 100 atoms of magnesium and divide by 100. (78.99)(23.9850) + (10.00)(24.9858) + (11.01)(25.9826) 100 60 POGILTM Activities for High School Chemistry

Average Atomic Mass - Mr. Sault's Classroom

Use one of the methods in Model 3 that gave the correct answer for average atomic mass to calculate the average atomic mass for oxygen. Isotope information is provided below. Show all of your work and check your answer against the mass listed on the periodic table. Isoto e Natural Abundance on Earth (0/0) Atomic Mass (am u) - 16.00 160 170 180

Pogil Answer Key Average Atomic Mass

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For the methods in Model 3 that gave the correct answer for average atomic mass, show that they are mathematically equivalent methods. Answer _ (78.99)(23.9850 amu) + (10.00)(24.9858 amu) + (11.01)(25.9826 amu)/ 100, (78.99)(23.9850 amu)/100 + (10.00)(24.9858 amu)/100 + (11..01)(25.9826 amu)/100 _____ 15.

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a. Calculate the average atomic mass of magnesium using data from Model 2. Hint: You will not get the correct answer if you add 24, 25 and 2Q and divide by 3. b. Give two reasons why your calculated value in part a is only an estimate of the average atomic mass of the element magnesium. ooby IQ's 10. The table below provides mass number and percent abundance information for the element lead.

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Describe the five parts of his atomic theory. JJ Thomson What was his occupation? Draw Thomson's apparatus What did Thompson do to cause the beam (cathode ray) to bend? What did he deduce from those measurements? What was the lightest known element? What was the new particle and

how much lighter was it? What model of the atom did Thomson propose?