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Introduction. In this particular lab we used stoichiometry, the part of chemistry that studies amounts of substances that are involved in reactions, to observe the reactions made by combining...

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Experiment 6 The Stoichiometry of a Reaction: The Molarity of a Solution
OBJECTIVES: To explore the stoichiometry of a chemical equation To perform a gravimetric analysis To run a double replacement reaction To work with solution molarities To determine the molarity of an unknown solution.
INTRODUCTION:

Lecture Notes 6 + Experiment 6 : STOICHIOMETRY OF ...

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stoichiometry, so during this lab, we got to put everything we have been learning to test and try it out The purpose of stoichiometry is to help scientists find out the moles of substance they need to use and how much of the product they will get out of it Experiment 6 ...

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to test and try it out. The purpose of stoichiometry is to help scientists find out the moles of substance they need to use and how much of the product they will get out of it.

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Stoichiometry of Reactions Kamie West Experiment 6 Section 14 October 21, 2015 David Starcevich I. Introduction This experiment will determine the amount of oxidizing agent, sodium hypochlorite (NaOCl), in bleach through a redox reaction. A known amount of bleach is mixed with water and this solution is acidified.

Lab 6 - Stoichiometry of Reactions Experiment 6 I Kamie ...

In this lab, we mixed together the reactants, 0.05 moles of baking soda and some vinegar into a flask. The products were the carbon dioxide, water, and sodium acetate. After mixing these chemicals...

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Pre-laboratory Assignment: Mole Ratios and Reaction Stoichiometry. Write balanced equations for the two reactions you will perform in this lab. Reaction \ref{3}: Reaction \ref{4}: Your goal in this lab is to experimentally verify the mole-to-mole ratios between a certain reactant and a certain product in both reactions.

7: Mole Ratios and Reaction Stoichiometry (Experiment ...

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Lab Experiment Stoichiometry of a Precipitation Reaction ...

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change in temperature (exothermic or endothermic) on the report sheet(6). Add 30 mL of 6.0 M NaOH to the solution in your beaker and describe the reaction of the report sheet (7). Add two or three boiling chips and carefully heat the solution – while stirring with a stirring rod – just to the boiling point.

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Eleventh grade Lesson

Stoichiometry Experimental Design

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