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Molarity \u0026amp; pH

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 $0.562 \text{ L} \cdot 10.78.9 \text{ g of CuSO}_4 \cdot 8H_2O$ is dissolved in 500.0 mL of water, calculate the molarity.
 $78.9 \text{ g} \times 1 \text{ mole} = 303.76 \text{ g} = 0.519 \text{ M}$
 0.5000 L
 Stoichiometry Worksheet # 3
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Molarity = moles / liters = $0.20 \text{ mole} / 1.5 \text{ L} = 0.133 \text{ M}$
 HNO_3 solution
 4. What volume (how many liters) of a 0.954 M solution can be made using 1.555 moles of H_2SO_4 ?
 liters = moles / Molarity = $1.555 \text{ moles} / 0.954 \text{ M} = 1.63 \text{ liters}$ of H_2SO_4 solution
 PART 2 5.

Molarity Wkst Answers - Worksheet Molarity ANSWER KEY PART ...

Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution.
 $1.54 \text{ mol } K_2SO_4 = 12.3 \text{ M } K_2SO_4$
 0.125 L soln . A chalice contains 36.45 grams ammonium chlorite in...

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What is the molarity of the following solutions given that: 1) 1.0 moles of potassium fluoride is dissolved to make 0.10 L of solution. $1.0 \text{ mole KF} = 10. \text{ M}$ 0.10 L soln 2) 1.0 grams of potassium fluoride is dissolved to make 0.10 L of solution. $1.0 \text{ g KF} \times 1 \text{ mole KF} = 0.0172 \text{ mol KF}$ 58 g KF $0.0172 \text{ mol KF} = 0.17 \text{ M}$ 0.10 L soln

Molarity Worksheet W 331 - Everett Community College

6. $375 \text{ mL} \times 0.0750 = 28.125 \text{ mL}$ ethylene glycol. $28.125 \text{ mL ethylene glycol} \times 1.09 \text{ g ethylene glycol/mL} = 30.7 \text{ g ethylene glycol}$. 7. $39 \text{ g KOH} \times 1 \text{ mole KOH} \times 1 \text{ L KOH} = 0.93 \text{ L} = 930 \text{ mL}$. $56 \text{ g KOH} \times 0.75 \text{ mol KOH}$. 8. $3.0 \text{ L soln} \times 0.750 \text{ moles HCl} \times 36.45 \text{ g HCl} = 82 \text{ g HCl}$. $1 \text{ L soln} \times 1 \text{ mole HCl}$.

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Solution: $MV = \text{grams} / \text{molar mass}$. $(x) (1.000 \text{ L}) = 245.0 \text{ g} / 98.0768 \text{ g mol}^{-1}$. $x = 2.49804235 \text{ M}$. to four sig figs, 2.498 M . If the volume had been specified as 1.00 L (as it often is in problems like this), the answer would have been 2.50 M , NOT 2.5 M .

ChemTeam: Molarity Problems #1 - 10

Calculate molarity of 35.0 mL KOH solution needed to completely neutralize 22.5 mL of $1.75 \text{ M H}_2\text{SO}_4$. Calculate volume (mL) of $2.50 \text{ M H}_2\text{SO}_4$ needed to completely neutralize 10.0 g NaOH (s) .
Answers. $M_1 V_1 = M_2 V_2$ $(1.71 \text{ M})(25.0 \text{ mL}) = M_2 (65.0 \text{ mL})$ $M_2 = 0.658 \text{ M}$; $M = \text{mol/L} = (25.0/40.0) / (0.325) = 1.92 \text{ mol/L}$

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moles KCl 55.2 g 74.6 g/mol Molarity Ph
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