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The first step is to replace the molecules in the square brackets with the units of concentration. Using our previous example: Kc = [HI]2 / [H2] [I2] = (mol dm-3)2 / (mol dm-3)(mol dm-3) You can see that the numerator, (mol dm-3)2, is the same as the denominator, (mol dm-3)(mol dm-3) and so all the units cancel.

How do you calculate the equilibrium constant, Kc, of a ...

Calculate Kc at this temperature. CH3COOH(l) + CH3CH2OH(l)

CH3COOCH2CH3(l) + H2O(l) .1.00 a-. 00. ..v.s.s.s o. XY<S. +0. (4) (A2 BBoF 29) 1.90 mol of hydrogen and 1.90 mol of iodine were allowed to reach equilibrium at 710 K. The equilibrium mixture contained 3.00 mol of HI.

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3 of 0.100 mol dm⁻³ sodium butanoate. c) 1.00 g of potassium
ethanoate is dissolved in 50 cm³ of 0.200 mol dm⁻³ ethanoic acid
(K_a = 1.74 × 10⁻⁵ mol dm⁻³). 2) Calculate the pH of the
following buffer solutions made by mixing an excess of weak

acids with strong bases. a) 25 cm³ of 0.5 mol dm⁻³ methanoic acid
(K_a = 1.6 × 10⁻⁴ mol dm⁻³).

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