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5.266 moles of TiCl_4 needs 10.53 moles of Mg to react, $\therefore \text{TiCl}_4$ is in XS and does not all react, so Mg is the limiting reagent $\therefore$ 2.058 moles of TiCl_4 reacts with 4.115 moles of Mg $\therefore$ 2.058 moles of Ti is produced
 Mass of Ti = $2.058 \times 47.9 = 98.6 \text{ g}$. © www.CHEMSHEETS.co.uk
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 $= 40.3 \times 0.741 / 40.3 = 0.0184$ 1 1 If used
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 consequential M2 (0.0185) 0.018 with no
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1.2 Amount Of Substance - A-Level Chemistry

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$\text{amount} = \text{mass}/M_r = 5 / (23.1 \times 2 + 12 + 16 \times 3) = 0.0472 \text{ mol conc=}$
 $\text{amount}/\text{Volume} = 0.0472 / 0.25 = 0.189 \text{ mol dm}^{-3}$
 Example 1: What is the amount, in mol, in 35.0g of CuSO_4 ?
 $\text{amount} = \text{mass}/M_r = 35 / (63.5 + 32 + 16 \times 4) = 0.219 \text{ mol}$
 It is usually best to give your answers to 3sf
 Example 3 : What is the volume in dm^3 at room

5. Formulae, equations and amounts of substance

Welcome to Topic 2 - AMOUNT OF SUBSTANCE. Topic 2 specification content
 Topic 2 notes

Topic 2 - Amount of Substance - A-Level Chemistry

In chemistry, the amount of substance in a given sample of matter is defined as the number of discrete atomic-scale particles in it divided by the Avogadro constant N_A . In a truly atomistic view, the amount of substance is simply the number of particles that constitute the substance. The particles or entities may be molecules, atoms, ions, electrons, or other, depending on the context. The value of the Avogadro constant N_A has been defined as $6.02214076 \times 10^{23} \text{ mol}^{-1}$. In the truly atomistic ...