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Dynamic Equilibrium At what point is a reversible reaction "completed?" Most of the reactions we have studied year have been "forward" reactions—once the reactant has been changed into the product it stays that way. The reaction "stops" when the limiting reactant runs out.

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Increase (T) causes the equilibrium to shift to the right (+) Decrease (I) causes the equilibrium to shift to the left Increase (T) causes the equilibrium to shift to the left Decrease (I) causes the equilibrium to shift to the right Temperature: A change in temperature corresponds to a change in energy therefore by using

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