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between atoms is called a covalent
bond, and the two electrons that join
atoms in a covalent bond are called a
bonding pair of electrons. A discrete

group of atoms connected by covalent
bonds is called a

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Metallic Bonds and Metallic Properties
When a metal is subjected to pressure,
the metal cations easily slide past one
another. If an ionic crystal is struck with
a hammer, the blow tends to push the
positive ions close together.

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When the elements react bonds are formed. Types of Bonds Ionic Metallic Covalent Ionic Bonds Bonds that are formed by transfer of electrons from one element to the other. Each element (now an ion) will have a complete octet after the transfer of electrons. The Ionic Bond

- The electrical force of attraction between oppositely charged ions.

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There are three major types of chemical bonding. They are the ionic bonding, covalent bonding and metallic bonding. The key difference between ionic bonding and metallic bonding is that the

ionic bonding takes place between positive and negative ions whereas the metallic bonding takes place between positive ions and electrons. Reference: 1. Libretexts. "Ionic and Covalent Bonds." Chemistry LibreTexts, National Science Foundation, 28 Feb. 2018. Available here 2.

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Lesson 7.1 Reading and Study Workbook

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Ionic and Metallic Bonding. BONDING AND INTERACTIONS. 7.1 Ions.

Essential Understanding Ions form when atoms gain or lose valence electrons, becoming electrically charged. Lesson Summary. Valence Electrons Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation. For a representative element, the group number equals the ...

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Group . 1A : 2A . 3A : 4A . 5A : 6A . 7A : 8A . 1 : 2 . 3 : 4 • Notice that all the electrons within a given group (with the exception of helium) have the same ...

Ionic and Metallic Bonding

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SECTION 7.2 IONIC BONDS AND IONIC

COMPOUNDS (pages 194–199) This section lists the characteristics of an ionic bond. It also describes the use of these characteristics to explain the electrical conductivity of ionic compounds when melted and when in aqueous solutions. Formation of Ionic Compounds (pages 194–195) 1. What is an ionic bond? 2.

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____ 12. The coordination number of an ion is the number of ions of positive charge that surround the ion in a crystal.
____ 13. The coordination number of the ion Na^+ in NaCl is 6. ____ 14. In forming an ionic compound, an atom of an element gains electrons. ____ 15.

05 CTR ch07 7/9/04 3:27 PM Page 155 IONS 7

Chapter 8: Covalent Bonding Pearson Chemistry Flashcards ... A covalent bond in which three pairs of electrons are shared by two atoms. polyatomic ion. A group of covalently bonded atoms ...

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Chapter 4 – Atoms and Covalent Bonds.
Chapter 5 – Covalent Bonds and
Molecular Structure. Chapter 6 –
Chemical ...

**Pearson Chemistry Covalent
Bonding Chapter Review Answers**

metallic crystal structures. Metallic
Bonds and Metallic Properties (pages
201–202) 1. Is the following sentence
true or false? Metals are made up of
cations, not neutral atoms. ____ 2. What

are metallic bonds? 3. Name three
properties of metals that can be
explained by metallic bonding. a. b. c. 4.
What happens to an ionic crystal when a
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