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Solution Example 1

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I. Determine what is oxidized and what is
reduced in each reaction. Identify the
oxidizing agent and the reducing agent,
also. 1. $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr 0 to Sr^{2+} ;
oxidized/reducing agent O_0 to O_2^- ;
reduced/ox. ag. 2. $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li 0 to
 Li^+ ; oxidized/red. ag. S_0 to S^{2-} ;
reduced/ox. ag. 3.

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Chapter 20 Worksheet: Redox I.
Determine what is oxidized and what is

reduced in each reaction. Identify the oxidizing agent and the reducing agent, also. 1.

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Chapter 20 Worksheet: Redox ANSWERS
 I. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also. 1. 2. 3. 4. 5. 6. $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ 2Li + S $\rightarrow$ Li₂S 2Cs + Br₂ $\rightarrow$ 2CsBr Sr() to Sr" ; oxidized/reducing agent 00 to ()2-: reduced/ox. ag. Li to Li oxidized/red, ag. so to reduced/ox. ag. NO to N₃"; reduced/ox. ag. (.)0 to (l) l w; reduced/ox. ag. Br to Br oxidized/red. ao.

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 reduced/ox. ag. 2. $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li 0 to Li^+ ; oxidized/red. ag. S^0 to S^{2-} ;
 reduced/ox. ag. 3. $2\text{Cs} + \text{Br}_2 \rightarrow 2\text{CsBr}$

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