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Chapter 7: Chemical Equilibrium 7.5-2
Statement: The value of Q is 0.867. Q is approximately equal to K , so the reaction is at equilibrium. 3. (a) The equilibrium constant equation for the reaction represented by $! 2 Ag 2O(s) \rightleftharpoons 4 Ag(s) + O_2(g)$ is $K = [O_2(g)]$, because the other substances are solids.

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